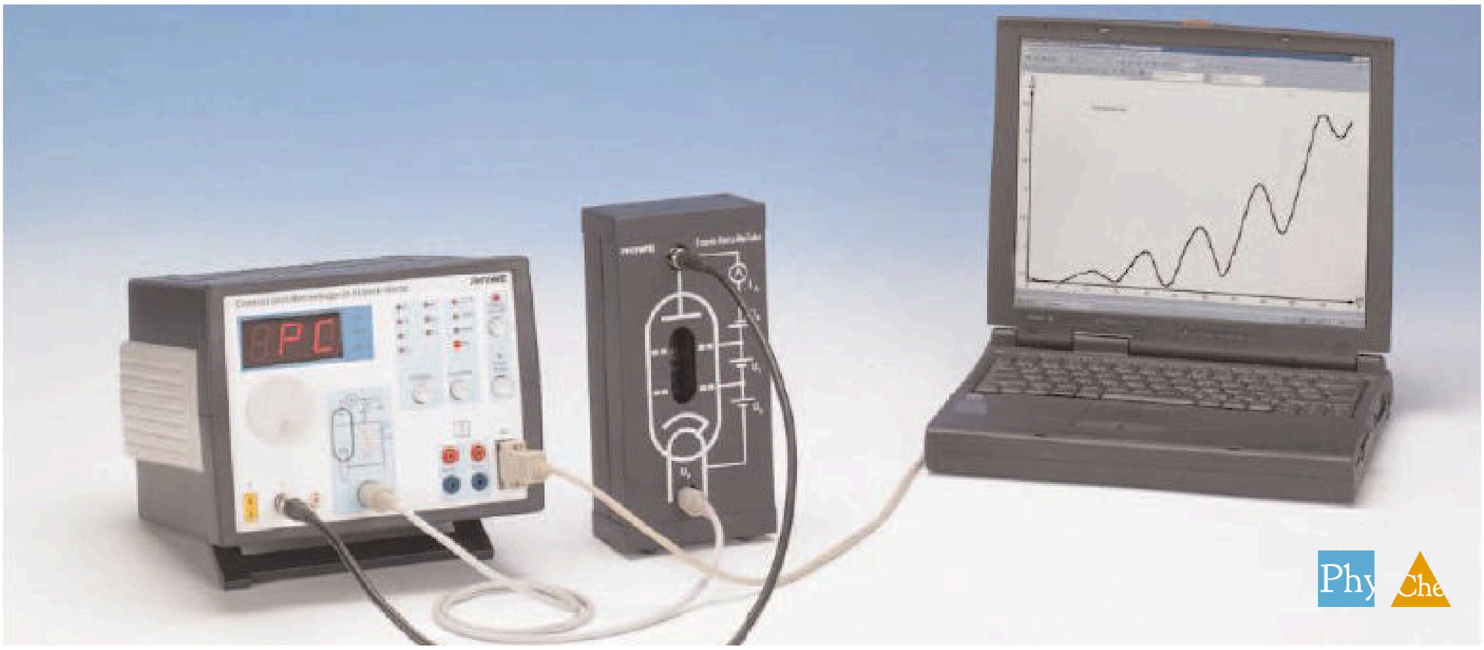
Franck-Hertz experiment with a Ne-tube



P2510315

Physics → Modern Physics → Quantum physics

Chemistry → Physical chemistry → Atomic structures & properties

Difficulty level	Group size	Preparation time	Execution time
easy	1	10 minutes	10 minutes

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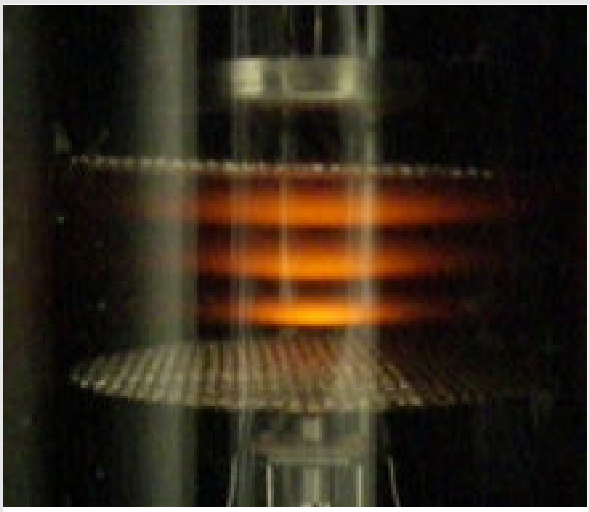
General information

Application



Franck-Hertz experiment with neon gas

The Franck-Hertz experiment shows the absorption of kinetic energy of electrons by neon atoms, or by mercury atoms in the original experiment.

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Other information (1/2)

Prior knowledge

The prior knowlege required for this experiment is found in the theory section.

Scientific principle

Electrons are accelerated in a tube filled with neon gas. The excitation energy of neon is determined from the distance between the equidistant minima of the electron current in a variable opposing electric field.

Other information (2/2)

Learning objective

Understanding energy quanta through Franck-Hertz experiment with neon gas

Tasks

1. Record the countercurrent strength I in a Franck-Hertz tube as a function of the anode voltage U .
2. Determine the excitation energy E from the positions of the current strength minima or maxima by difference formation.

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Safety instructions

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For this experiment the general instructions for safe experimentation in science lessons apply.

For H- and P-phrases please consult the safety data sheet of the respective chemical.

Theory (1/5)

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Niels Bohr introduced the planetary model of the atom in 1913: An isolated atom consists of a positively charged nucleus about which electrons are distributed in successive orbits. He also postulated that only those orbits occur for which the angular momentum of the electron is an integral multiple of $h/2\pi$, i.e. $n \cdot h/2\pi$, where n is an integer and h is Planck's constant.

Bohr's picture of electrons in discrete states with transitions among those states producing radiation whose frequency is determined by the energy differences between states can be derived from the quantum mechanics which re-placed classical mechanics when dealing with structures as small as atoms.

It seems reasonable from the Bohr model that just as electrons may make transitions down from allowed higher energy states to lower ones, they may be excited up into higher energy states by absorbing precisely the amount of energy representing difference between the lower and higher states.

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Theory (2/5)

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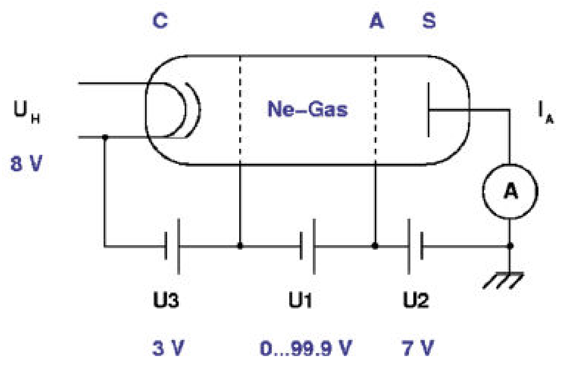
James Franck and Gustav Hertz showed that this was, indeed, the case in a series of experiments reported in 1913, the same year that Bohr presented his model. Franck and Hertz used a beam of accelerated electrons to measure the energy required to lift electrons in the ground state of a gas of mercury atoms to the first excited state.

In present experiment, a tube filled with neon gas is used. The electrons emitted by a thermionic cathode are accelerated between cathode C and anode A in the tube filled with neon gas and are scattered by elastic collision with neon atoms.



Principle of the measurement.

Theory (3/5)

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From an anode voltage U_1 of 16,8 V, however, the kinetic energy of the electrons is sufficient to bring the valence electron of the neon to the first excitation level by an inelastic collision. Because of the accompanying loss of energy, the electron can now no longer traverse the opposing field between anode A and counter electrode S: the current I is at a minimum.

If we now increase the anode voltage further, the kinetic energy of the electron is again sufficient to surmount the opposing field: the current strength I increases.

When U_1 = 2 · 16.8 V the kinetic energy is so high that two atoms in succession can be excited by the same electron: we obtain a second minimum.

These minima are not, however, very well-defined because of the initial thermal distribution of the electron velocities.

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Theory (4/5)

The voltage U_1 between anode and cathode is represented by

$$U_1 = U + (\Phi_A - \Phi_C)$$

where U is the applied voltage, and Φ the work function voltages of the anode and cathode re-spectively. As the excitation energy E is determined from the voltage differences at the minima, the work function voltages are of no significance here.

According to the classical theory the energy levels to which the mercury atoms are excited could be random. According to the quantum theory, however, a definite energy level must suddenly be assigned to the atom in an elementary process.

The course of the I/U_A curve was first explained on the basis of this view and thus represents a confirmation of the quantum theory. The excited neon atom again releases the energy it has absorbed, with the emission of a photon.

Theory (5/5)

When the excitation energy E is 16.8 eV, the wavelength of this photon is

$$\lambda = \frac{ch}{E}$$

where $c = 2.9979 \cdot 10^8 \ \frac{m}{s}$ and $h = 4.136 \cdot 10^{-15} \ eVs$

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Equipment

Position	Material	Item No.	Quantity
1	Franck-Hertz Ne-tube w. housing	09105-40	1
2	Franck-Hertz control unit	09105-99	1
3	Connect.cord f.Franck-H. Ne-tube	09105-50	1
4	Software Measure Franck-Hertz experiment	14522-61	1
5	Converter USB - RS232, active	14602-10	1
6	Data cable, plug/ socket, 9 pole	14602-00	1
7	Screened cable, BNC, l = 750 mm	07542-11	1

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Additional equipment

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Position	Material	Quantity
1	PC	1

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Setup and procedure

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Setup

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Experimental setup

Set up the experiment as shown in the figure. For details see the operating instructions of the unit 09105-99.

Connect the Franck-Hertz operating unit to the computer port COM1, COM2 or to USB port (use USB to RS232 Adapter Converter 14602-10).

Procedure

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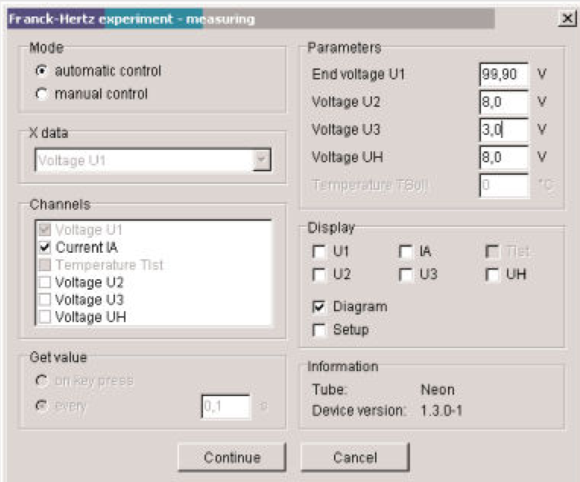
Start the measure program and select Franck- Hertz experiment Gauge. The window “Frank- Hertz-experiment – measuring” appears. The optimum parameters are different for each Ne-tube. You find the specific parameters for your device on a sheet which is enclosed in the package of the Ne-tube.

Choose the parameters for U1, U2, U3 and UH as given on that sheet and make sure that the rest is set as shown in Fig. 2. Press the continue button.



Measuring parameters.

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Evaluation (1/2)

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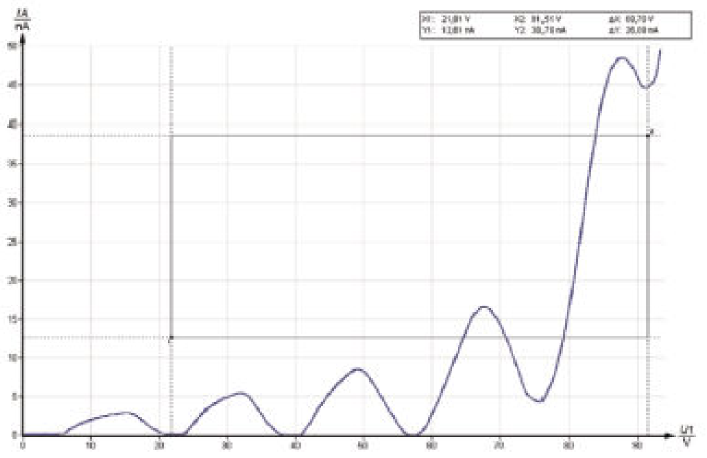

Example of a Franck-Hertz curve recorded with Ne-tube.

The graph of I/U_1 shows equidistant maxima and minima.

For our evaluation we determine the voltage values of the minima. From the differences between these values we obtain the excitation energy E of the neon atom by taking an average. By evaluating the measurements in the figure we obtained the value

$$E = (17.4 \pm 0.7)\, eV$$

Evaluation (2/2)

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Fill in the blanks:

In a tube filled with ________, electrons are emitted by glow emission from a thermionic ________ and accelerated to the ________ by the acceleration voltage. If the voltage is increased slowly, the measured current values initially ________ exponentially, up to a certain voltage and then ________ slowly. The current values ________ again when voltage is twice the value of the initial voltage (at which the current rise for the first time).

- anode
- increase
- cathode
- decrease
- increase
- neon gas

Check

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Slide	Score / Total
Slide 17: Summary experiment	0/6

Total Score 0 / 6

Show solutions   Retry

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